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Figure 9.1. 1: Gas Pressure. Pressure is what results

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when gas particles rebound off the walls of their container. The basic unit of pressure is the newton per square meter ( $\text{N/m}^2$ ). This combined unit is redefined as a pascal (Pa).

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1. all gases have mass  
2. particles in a gas are separated from each other by large distances  
3. gases are always moving in random motion  
4. gases do not lose energy when they collide because they are "perfectly elastic"  
5. gases have no attraction or IMF's  
6. gases exert pressure on the walls of their container

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Chemistry – Unit 9. Worksheet 1 – Gases Again. Use gas laws to solve # 1 –3, doing the work like we did in Unit 2 (PVT chart). Use new information to solve # 4 –8. 1. A can of spray paint contains nitrogen gas as the propellant. The pressure of the gas is 3.5 atm



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when the temperature is  $20.^{\circ}\text{C}$ .

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In a mixture of gases the total pressure of the mixture is equal to the sum of the pressures that each gas

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would exert by itself in the same volume. You have 1 L of oxygen gas at a pressure of 1 atm, 1 L of nitrogen gas at a pressure of 2 atm, and 1 L of hydrogen gas at a pressure of 3 atm.

### Chemistry Unit 9 Flashcards | Quizlet

A gas is a state of matter with no defined shape or volume. Gases have their own unique behavior depending on a variety of variables, such as temperature, pressure, and volume. While each gas is different, all gases act in a similar matter. This study guide highlights the concepts and laws dealing with the chemistry of gases.

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Chemistry -Unit 9 -Worksheet 1 -Gases & Mixtures.

Complete the following problems on a SEPARATE  
PIECE OF PAPER. Show all work in addressing the  
problem, including any modeling / charts Box your  
answer with complete labels! 1. A can of spray paint  
contains nitrogen gas as the propellant.

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9.1 Gas Pressure; 9.2 Relating Pressure, Volume,

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Amount, and Temperature: ... These conditions will decrease the both the frequency of molecule-wall collisions and the number of collisions per unit area, the combined effects of which balance the effect of increased collision forces due to the greater kinetic energy at the higher temperature ...

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33. A piston compresses a gas at constant temperature. Initially the gas occupied 1 L and was at a pressure of 1 atm. After compression, the gas occupies 0.1 atm. What is the pressure of the compressed gas? 0.1 atm 1 atm 10 atm 100 atm



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